Easy Functionalization of Carbon Nano-onions with Disulfides and their use as Recyclable Heterogeneous Organocatalysts

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Abstract

A facile functionalization of CNOs, through an atom economy one-step process, with a series of different generation (2.0, 2.5, and 3.0G) poly(amidoamine) (PAMAM) dendrimers with a cystamine core and endowed with either methyl ester or amino groups in the periphery has been reported. The radical addition reaction onto CNOs surface, after the homolytic rupture of the S–S bond of cystamine, afforded CNO-PAMAM hybrids, which were extensively characterized by the use of various analytical and spectroscopic techniques such as TGA, FT-IR, solid state ^{13}C-NMR, Raman, XPS, DLS and zeta potential. Furthermore, the presence and availability of -NH_2 groups has been demonstrated by applying the functionalized CNOs with PAMAM 2.0G and 3.0G as heterogeneous and recyclable organocatalysts in the one-pot three-component synthesis of 2-amino-4*H*-chromene derivatives.

Keywords: Carbon nano onions; PAMAM dendrimers; Heterogeneous catalysis; 2-Amino-4H-chromene derivatives; Three-component one-pot synthesis.

1. Introduction

The ability of carbon to exist in several nanoallotropic forms comprising 0D, 1D, and 2D structures organized in strings, tubes, sheets, spheres, or multi-layered spheres (nano-onions), has attracted tremendous interest from the scientific community due to their potential use in several applications, mainly in the field of nanotechnology [1-5]. Despite their outstanding mechanical, physical, and electronic properties, carbon nanomaterials (CNMs) usually display poor solubility and bad dispersibility in common solvents that strongly hampers their wider use, limiting their potential.

Carbon nano-onions (CNOs) structurally consist of multiple concentric fullerene-like shells. As is typical for a member of the CNM family, CNOs show aggregation due to the strong intermolecular electrostatic interactions, dominated by van der Waal forces, established through their large external surface area.

To overcome the poor dispersibility of CNOs, their surface chemical functionalization may represent a double solution; on the one hand, preventing their aggregation and, on the other, allowing to obtain new materials that combine CNOs' intrinsic properties with those of the functionalizing moieties [6]. In this regard, protocols [7] have been adopted for covalent modification of CNOs: besides the classic oxidation and amidation strategy, which deeply alter the CNO-surface [8, 9], other chemical functionalization includes fluorination [6], [2+1] cycloaddition of nitrenes [10], Prato reaction [9, 11], Tour reaction [12, 13] and more recently the Birch reaction [14, 15].

As stated, the chemical derivatization of CNO along with the possibility for their gram-scale synthesis has allowed to use the functionalized materials for several applications such as energy and

electrochemical storage [16, 17], sensing [18-20], supercapacitors [21, 22], fuel cells [23] and water treatment [24].

Significantly, the versatility [25] of CNOs has given way to their biomedical applications [26]. Their biocompatibility and easy penetration into cells, demonstrated through their use in bioimaging [27, 28], their efficient delivery of glycopeptides and proteins [29], and their targeted delivery of anticancer therapeutics [30], underline their capacity as non-cytotoxic carriers for cellular uptake [31].

The use of CNOs as catalysts has been less explored despite their promising roots in the oxidative dehydrogenation of ethylbenzene [32, 33]. Pristine or oxidized CNOs were employed as nanocomposites with Bi_2O_3 or NiCu nanoparticles for the photocatalytic thiol-ene reaction [34] or C-C coupling processes [35], respectively. Nitrogen-rich CNOs [36] and nitrogen-doped CNOs with molybdenum carbide [37] have been, in turn, used as heterogeneous catalysts for the oxygen reduction reaction and alkene dehydrogenation. Recently, a series of CNO-metallophthalocyanine conjugates have been applied as photocatalysts for degrading rhodamine B [38].

However, except for the recent case, the above studies deal only with unfunctionalized or oxidized CNOs. Moreover, to the best of our knowledge, no reports have been made of the possible reuse of CNO-based catalysts. In this regard, we believe CNOs represent a viable scaffold for developing recyclable heterogeneous catalysts.

Nevertheless, new versatile and reliable methods for the covalent functionalization of CNOs as chemical tools for their easy manipulation are always welcome, especially those avoiding harsh conditions that may alter the electronic properties, since they allow accede to novel hybrid advanced functional materials.

Herein, we report the facile functionalization of CNOs, through an atom economy one-step process, with a series of poly(amidoamine) (PAMAM) dendrimers endowed with either methyl ester or amino

groups in the periphery, and a cystamine core of varying generation. The presence and availability of -NH_2 groups has been demonstrated by applying the corresponding hybrids as heterogeneous and recyclable organocatalysts in the one-pot three-component synthesis of 2-amino-4*H*-chromene derivatives.

2. Materials and methods

Chemicals and solvents were purchased from commercial suppliers and used as received without further purification. Cystamine-core PAMAM dendrimers were prepared according to the reported literature procedures [39, 40].

Thermogravimetric analysis (TGA) was performed under nitrogen flow from 100 to 1000 °C with a heating rate of 10 °C min^{-1} with a Mettler Toledo TGA/DSC STAR System. The temperature was increased starting from RT up to 100 °C, and then this temperature was kept for 30 min to remove adsorbed water or volatile solvents before reaching 1000 °C.

^{13}C CP-MAS NMR spectra were acquired on a Bruker Advance II 400 spectrometer operating at 100.63 MHz for ^{13}C nuclides and 400.15 MHz for ^{1}H nuclides equipped with a 4 mm (H–X) double channel CPMAS probe. The samples were placed in a 4 mm zirconia rotor closed with Kel-F caps. The spectra were measured as 1024 scans at a MAS speed of 8 kHz, with 2 ms contact time, 3 s delay time, and an excitation pulse of 4.7 ms on the ^{1}H nucleus. The Hartmann–Hahn condition was optimized using an adamantane standard sample, which was also used as an external chemical shift reference.

ATR-FTIR spectra were recorded on a Thermo Scientific Nicolet Summit Spectrometer with an Everest Diamond ATR accessory over the range of 4000-500 cm^{-1}. The resolution was set to 4 cm^{-1} and the velocity of the mirror was set to 0.4747 cm/s, with background measurement was performed for every sample. The spectra of functionalized CNOs were taken at 128 scans to improve S/N ratio, with all other samples taken at 32 scans. Detection was carried out by a DTGS KBr detector with the

aperture set to large (100). Prior to FT, apodization was done using the Norton Beer Strong method to mitigate truncation errors. A zero-filling factor of 2 was used to avoid clipping of spectral bands, and automatic atmospheric suppression was enabled.

XPS analyses were performed with a VGMicrotech ESCA 3000Multilab, equipped with a dual Mg/Al anode. The spectra were excited with the unmonochromatized Al Kα source (1486.6 eV) run at 14 kV and 15 mA. The analyzer was operated in the constant analyzer energy mode. A pass energy of 20 eV set across the hemispheres was used for the individual peak energy regions. Survey spectra were measured at 50 eV pass energy. The sample powders were mounted on a double - sided adhesive tape. The pressure in the analysis chamber was in the range of 10^{-8} Torr during data collection. The constant charging of the samples was removed by referencing all the energies to the C 1*s* set at 284.6 eV. The invariance of the peak shapes and widths at the beginning and end of the analyses ensured the absence of differential charging. Analyses of the peaks were carried out with the software provided by VG, based on the non-linear least-squares fitting program using a weighted sum of Lorentzian and Gaussian component curves after background subtraction, according to Shirley and Sherwood [41]. Atomic concentrations were calculated from peak intensity using the sensitivity factors provided by the software. The binding energy values are quoted with a precision of ± 0.15 eV, and the atomic percentage with a precision of ± 10%.

DLS and ZP studies were carried out on CNOs functionalized with three generations of PAMAM; namely, **CNO-PAMAM 2.0G**, **CNO-PAMAM 2.5G**, and **CNO-PAMAM 3.0G**. Each material was analysed at 5, 10 and 20 ppm; briefly, all samples were dispersed in deionized H_2O and sonicated for 15 min prior to analysis. The sonic-bath temperature was maintained at 22-26 °C, and samples were equilibrated to 25 °C prior to DLS, and subsequent ZP, analysis. The analyses were carried out on a ZetaSizer Ultra instrument in backscatter mode, using a PCS1115 cell for DLS measurements and a DTS1070 cell for ZP measurements. Each sample was analyzed 3-6 times each by DLS and ZP, with measurements carried out in 5 min intervals. ZS XPLORER software was used for instrument control,

data acquisition and statistical analysis. All standard deviations are calculated from multiple ($n \geq 3$) experiments, and mean diameters are reported from the highest intensity peaks. Data was plotted using a custom python script utilizing the matplotlib package.

2.1 Preparation of CNO-PAMAM 2.0G

In a 50 mL flask, 15 mg of p-CNOs (1.25 mmol) were suspended in 10 mL of toluene and sonicated for 15 min. A solution of 1.0 g PAMAM 2.0G (0.66 mmol) in 4 mL of methanol was added to the p-CNOs suspension and degassed for 15 min. Then, 25 mg of AIBN (0.15 mmol) were added and the mixture was heated at 100 °C for the indicated time under Ar atmosphere. The resulting dispersion was allowed to cool, filtered on a Millipore membrane (PTFE 0.45 μm) and the residue was washed with toluene, MeOH, water, 3.0 M HCl, 2% v/v triethylamine in THF, THF and diethyl ether. The resulting black solid was dried under vacuum at 40 °C overnight.

2.2 Preparation of CNO-PAMAM 2.5G(I and II)

In a 50 mL flask 15 mg of p-CNOs (1.25 mmol) were suspended in 8 mL of toluene and sonicated for 15 min. A solution of 1.0 g PAMAM 2.5G (0.35 mmol) in 2 mL of methanol was added to the p-CNOs suspension and the mixture was heated at 100 °C for 72 h, in the case of **CNO-PAMAM 2.5G(I),** or 96 h, in the case of **CNO-PAMAM 2.5G(II)**. The resulting dispersion was allowed to cool, filtered on a Millipore membrane (PTFE 0.45 μm), and the residue was washed with toluene, MeOH, and diethyl ether. The resulting black solid was dried under vacuum at 40 °C overnight.

2.3 Preparation of CNO-PAMAM 3.0G(I and II)

In a 50 mL flask 15 mg of p-CNOs (1.25 mmol) were suspended in 10 mL of toluene and sonicated for 15 min. A solution of 1.0 g PAMAM 3.0G (0.30 mmol) in 4 mL of methanol was added to the p-CNOs suspension and degassed for 15 min. Then, 25 mg of AIBN (0.15 mmol) were added and the mixture was heated at 100 °C for 96 h, in the case of **CNO-PAMAM 3.0G(I),** or 120 h, in the case

of **CNO-PAMAM 3.0G(II),** under Ar atmosphere. The resulting dispersion was allowed to cool, filtered on a Millipore membrane (PTFE 0.45 μm) and the residue was washed with toluene, MeOH, water, 3.0 M HCl, 2% v/v triethylamine in THF, THF, and diethyl ether. The resulting black solid was dried under vacuum at 40 °C overnight.

2.4 Synthesis of 2-amino-3-cyano-4-(4-bromophenyl)-4H-benzo[h]chromene (1)

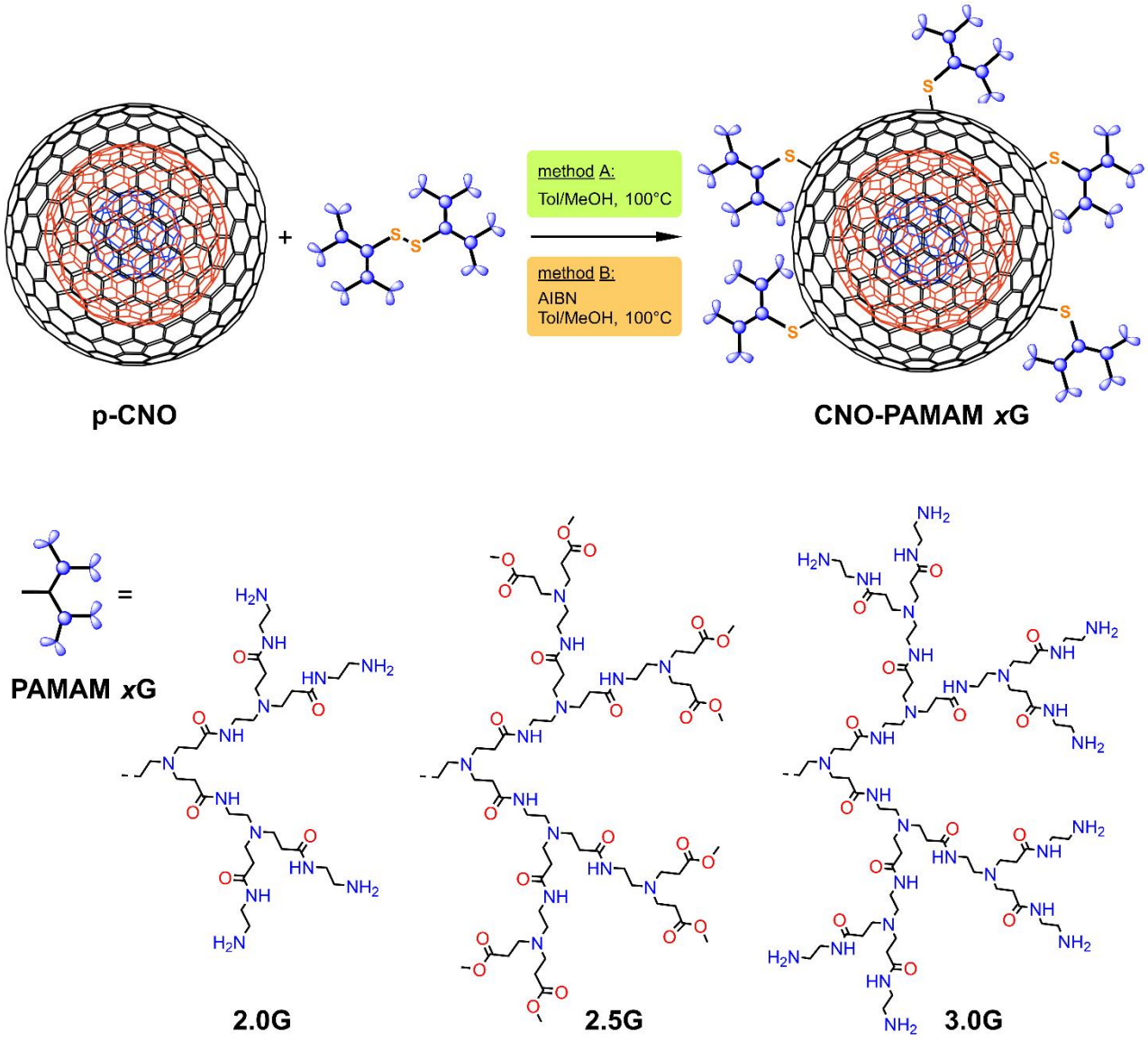
In a 5 mL glass vial with screw cap, 1-naphthol (0.2 mmol), malononitrile (0.2 mmol), 4-bromobenzaldehyde (0.2 mmol), and the catalyst (6 mg; ~3 mol% for **CNO-PAMAM 2.0G**; ~10 mol% **CNO-PAMAM 3.0G(II)**) were introduced and 0.5 mL deionized water was added. The reaction mixture was shortly sonicated to disperse catalyst and stirred at 70 °C for 2 h. The reaction mixture was allowed to cool down to room temperature. Afterwards acetone was added before sonication to dissolve all the reactants/product and recover the catalyst by centrifugation. The washings with acetone were repeated twice and all supernatants were combined, dried under reduced pressure, and analyzed by ^{1}H NMR spectroscopy. The recovered catalyst was dried at 60 °C before being reused in the next cycle.

2.5 Synthesis of 2-amino-4-(3-bromophenyl)-7,7-dimethyl-5-oxo-5,6,7,8-tetrahydro-4H-chromene-3-carbonitrile (2)

In a 5 mL glass vial with screw cap, dimedone (0.25 mmol), malononitrile (0.3 mmol), 4-bromobenzaldehyde (0.25 mmol), and the catalyst (4 mg, ~5 mol% of **CNO-PAMAM 3.0G(II)**) were introduced and 0.5 mL deionized water was added. The reaction mixture was shortly sonicated to disperse catalyst and stirred at 90 °C for 1 h. The work up procedure was the same as described above.

3. Results and Discussion

The hybrids **CNO−PAMAM 2.0G**, **2.5G**, and **3.0G** (**Scheme 1**) have been prepared through the atom-economical strategy which involves the direct reaction of pristine CNO (p-CNOs) with disulfides—namely, PAMAM dendrimers with *x*G cystamine core, where *x* is generation 2.0, 2.5 or 3.0—through a toluene-reflux protocol developed for other carbon nanoforms [42-44]. The full-generation dendrimer (2.0G or 3.0G) is comprised of two or three layers of branching points and has amino groups at the ends, whereas the middle generation (2.5G) indicated two layers and the presence of methyl ester group at the ends. PAMAM dendrimers have found application as macromolecular vectors for drug delivery [45-48] as well as in biomedical applications [49-51], and the presence of -NH_2 groups and methyl esters in the periphery allows for their possible post-functionalization through amidation, nucleophilic substitution or formation of carbamates for the former, and hydrolysis and esterification for the latter. To functionalize CNOs with the full generation PAMAM dendrimers (2.0G and 3.0G), azobisisobutyronitrile (AIBN) was utilized, as per analogous reactions between thiols and several other carbon nanomaterials (CNMs) [52-54].



Scheme 1. Synthesis of CNO-PAMAM *x*G hybrids from p-CNOs through methods A and B, and the structure of the corresponding PAMAM generations.

The reactions were initially kept for four days for the full generation PAMAM (2.0G and 3.0G) and three days for PAMAM 2.5G (**Table 1**, entries 1-3), but more functionalized materials were obtained when the reaction with PAMAM 2.5G and 3.0G were maintained one day longer (**Table 1**, entries 4-5; see Materials and Methods for conditions). For clarity, the reactions of entries 2 and 3 were left over the weekend, and in both cases, most of the dendrimer adhered to the flask walls above the solvent level. In the repeated reactions, we took care to push, each day of reaction, the dendrimer below the solvent level to be in contact with the solvent and stirred. It is worth mentioning that the

title process is 100% atom economy since all the atoms are linked to the CNO, and the cystamine-cored PAMAM that did not react can be readily recovered by filtration at the end of the process.

Table 1. Preparation of CNO-PAMAM hybrids and loading in PAMAM dendrons and amino groups.[a]

Entry	Material	Method	Solvent ratio (mL)	Time (h)	Loading (mmol/g)[b]	-NH_2 content (mmol/g)
1	**CNO-PAMAM 2.0G**	B	MeOH:Tol 4:10	96	0.235	0.94
2	**CNO-PAMAM 2.5G(I)**	A	MeOH:Tol 2:8	72	0.141	-----
3	**CNO-PAMAM 3.0G(I)**	B	MeOH:Tol 4:10	96	0.127	1.04
4	**CNO-PAMAM 2.5G(II)**	A	MeOH:Tol 4:16	96	0.484	-----
5	**CNO-PAMAM 3.0G(II)**	B	MeOH:Tol 4:10	120	0.437	3.48

[a] Reaction conditions: Method A – 15 mg of CNOs, 1 g of PAMAM *x*G in methanol/toluene mixture at 100 °C; method B – 15 mg of CNOs, 1 g of PAMAM *x*G, 25 mg of AIBN in methanol/toluene mixture at 100 °C. [b] Determined by TGA at 700 °C.

Materials **CNO-PAMAM 2.0G**, **CNO-PAMAM 2.5G(I** and **II)**, and **CNO-PAMAM 3.0G(I** and **II)**, as well as the parental p-CNOs and the three PAMAM generations, were subjected to thermogravimetric analysis (TGA) under nitrogen flow to study both their thermal stability and the functionalization degree of p-CNOs. The TGA curves and derivative thermogravimetry (DTG) curves have been presented in **Figure 1**, wherein the relevant wt% values have also been outlined. Given that no decomposition of p-CNOs was observed up to 700 °C, these wt% values were quantified at 700 °C, ensuring that all organic coatings were decomposed. Further estimation of the functionalization degree of the hybrids has been carried out to assess their PAMAM and peripheral amino content; these values are reported in **Table 1**.

For **CNO-PAMAM 2.0G**, a net weight loss of 15.8% can be observed, corresponding to a loading of 0.94 mmol/g of peripheral amino groups, considering the residual 11.7 wt% for PAMAM 2.0G

(**Figure 1a**). In addition, the PAMAM-modified CNOs show increased thermal stability compared to their parental PAMAM. **CNO-PAMAM 2.5G(I)** and **CNO-PAMAM 2.5G(II)** show similar thermal behavior with a weight loss of 20.5 wt% and 70.1 wt%, respectively (**Figure 1b**). This indicates that the reaction time can impart control over the degree of functionalization. TGA of **CNO-PAMAM 3.0G(I)** and **CNO-PAMAM 3.0G(II)** indicates that both hybrids are stable somewhat beyond 200 °C and a net weight loss of 21.2% and 73.2%, respectively, corresponding to 1.04 and 3.48 mmol/g of amino groups (**Figure 1c**).

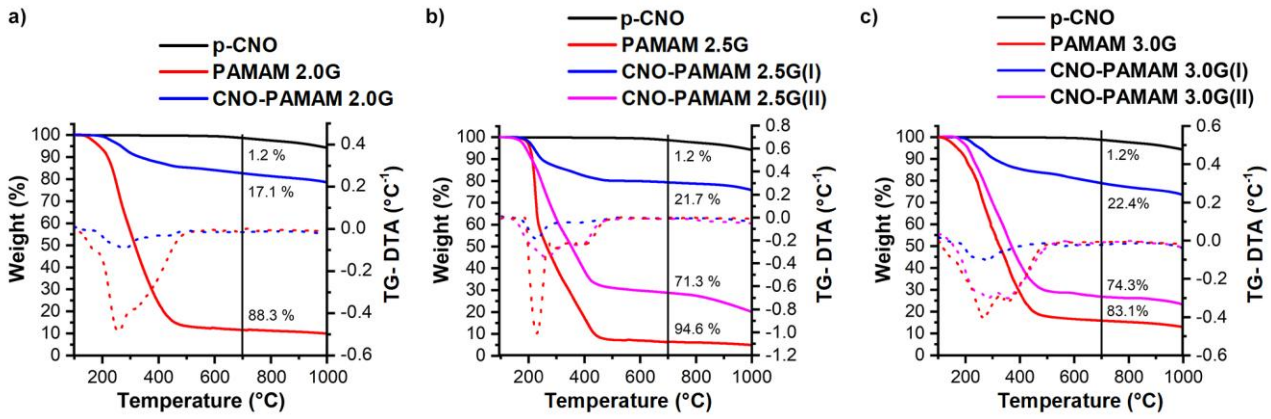
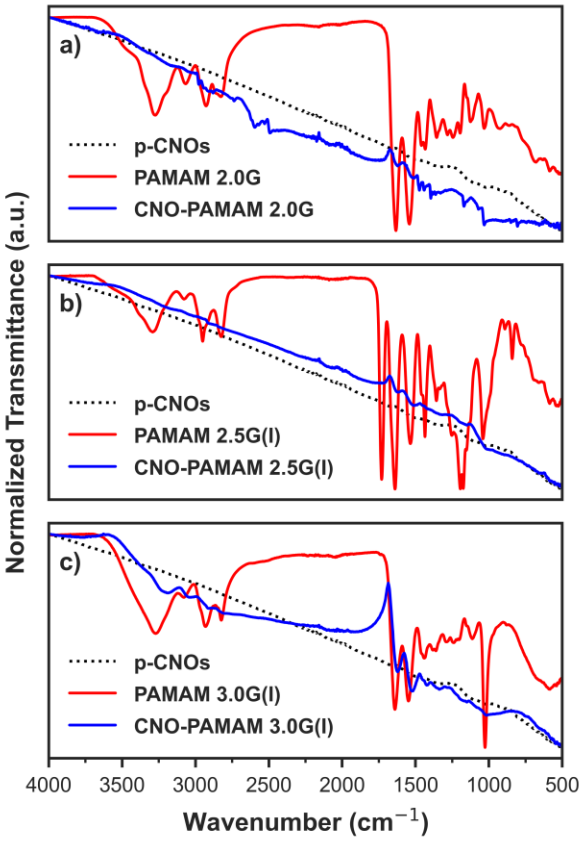


Figure 1. TGA (solid) and DTG (dashed) curves of p-CNOs, PAMAM *x*G dendrimers, and the CNO-PAMAM *x*G hybrids.

The ATR-FTIR spectra of all the prepared materials and those of the dendrimers used for their functionalization are reported in **Figure 2**. This simple and cheap technique allowed us to detect the functional groups present in our materials. As expected, in the spectra of **CNO-PAMAM 2.0G** and **CNO-PAMAM 3.0G(I)**, we find the presence of two signals at ~1640 and ~1540 cm^{-1} due to the stretching of the amide C=O and the bending of the amide N-H group, respectively, and, in the 3300-3200 cm^{-1} range weak signals are found due to the presence of amino groups.



1

2 **Figure 2.** ATR-FTIR spectra of p-CNOs (*dashed*), and the three generations of the PAMAM
3 dendrimer and the CNO-PAMAM *x*G hybrids (*solid*).

4 For the first time, it was possible to obtain solid-state ^{13}C NMR spectra of functionalized-CNO by
5 using the "Cross-Polarization Magic Angle Spinning" technique (CP-MAS ^{13}C NMR) for **CNO-**
6 **PAMAM 2.5G(II)** and for **CNO-PAMAM 3.0G(II)** (**Figure 3**).

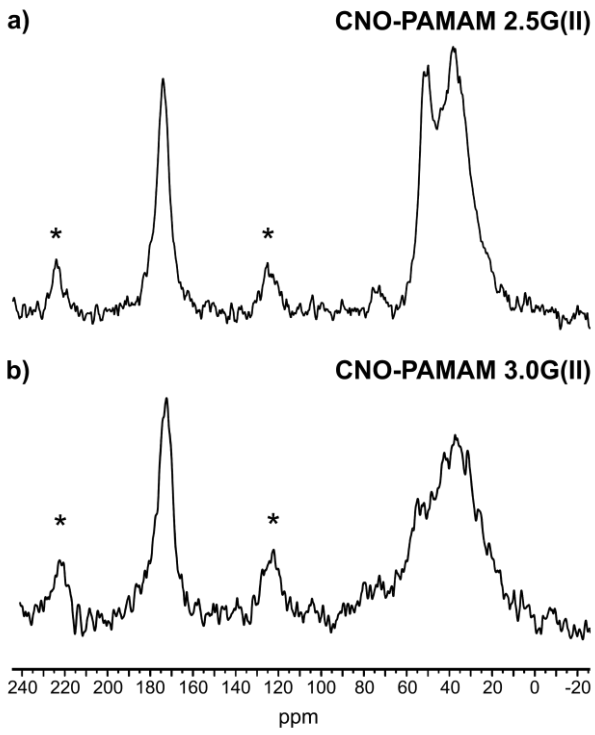


Figure 3. Solid-state ^{13}C-NMR spectra of (a) **CNO-PAMAM 2.5G(II)**; and (b) **CNO-PAMAM 3.0G(II)** hybrids. Asterisks indicate sidebands.

The ^{13}C-NMR spectrum of **CNO-PAMAM 2.5G** (**Figure 3a**) shows a broad signal comprised of two peaks centered at about 35 and 50 ppm. The first corresponds to carbon atoms in α to the amide nitrogen and to the –CH_2– in α to the carbonyl, and the second to the carbon atoms in α to the tertiary amino nitrogen and to the methoxy groups. A well-defined peak at 175 ppm is also observed, which belongs to the carbonyl groups of PAMAM, fully confirming the proposed structure.

The spectrum of **CNO-PAMAM 3.0G(II)** is represented in **Figure 3b** and, as in the previous case, it is possible to observe the peak at 174 ppm due to the carbonyl carbon atoms. The aliphatic region appears as a single large and complex signal between 20 and 60 ppm. This is due to the presence of numerous peripheral methylenes of ethylenediamine, which populate the area between 37 and 42 ppm. Again, the spectrum is in excellent agreement with the proposed structure, confirming the successful covalent functionalization of the CNOs.

Raman spectroscopy was carried out on p-CNOs and the three CNO-PAMAM *x*G hybrids. The spectra, as well as the calculated I_D/I_G and I_{2D}/I_G ratios, have been presented in (**Figure 4**).

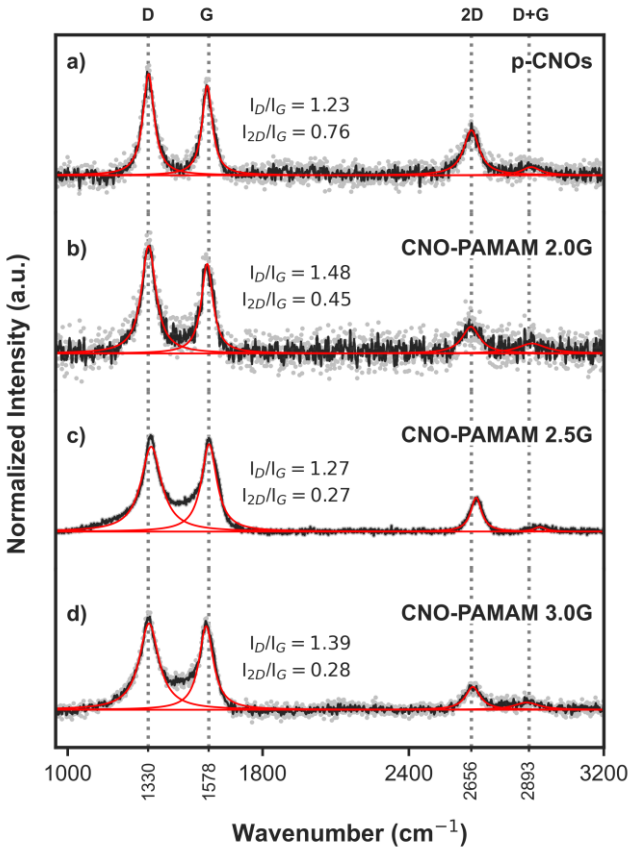


Figure 4. Raman spectra of p-CNOs and the three CNO-PAMAM *x*G hybrids. The spectra (*black*) represent an average of $n = 3$ scans (*gray, scatter*). The I_D/I_G and I_{2D}/I_G ratios were calculated from the intensity integral of deconvoluted bands (*red*).

The Raman spectroscopy data (**Figure 4**) reveals that PAMAM dendrimer functionalization of CNOs impacts both disorder (I_D/I_G ratio) and electronic properties (I_{2D}/I_G ratio) in nuanced ways. The I_D/I_G ratio increases most substantially from p-CNOs to **CNO-PAMAM 2.0G**, suggesting an increase in defects or disorder introduced by the initial dendrimer functionalization. Interestingly, the I_D/I_G ratio increases to a lesser extent the case of **CNO-PAMAM 3.0G(II)** and least so in the case of **CNO-PAMAM 2.5G(I)**. There are several contributing factors for which a relatively lower increase could be observed. The foremost reason is the degree of functionalization; a higher surface loading is proportional to the increase in the I_D/I_G ratio, as more sp^2 carbon is converted into sp^3. However, these two generations of dendrimers could selectively passivate surface defect sites, which might reduce the Raman spectroscopic signature associated with defects, leading to a lower perceived I_D/I_G ratio. Notwithstanding, all three CNO-PAMAM *x*G hybrids demonstrate an increase in the I_D/I_G ratio,

indicating successful surface covalent functionalization of the p-CNO surface in the preparation of the three CNO-PAMAM *x*G hybrids. This indication is further supported by a pronounced decrease in the I_{2D}/I_G ratio observed for all three hybrids when compared to their parental p-CNOs. Moreover, as dendrimer generation increases, a general trend of decreasing I_{2D}/I_G ratio is further observed. These trends suggest alterations in electronic structure and interlayer interactions of the materials [55]—notably, the decrease in the I_{2D}/I_G ratio aligns with a general trend towards thicker layered structures as dendrimer generation increases.

All the materials were characterized by X-ray photoelectron spectroscopy (XPS) to evaluate the presence, the oxidation state and the surface concentration of the elements contained, and this data is collected in **Table 2**. The high-resolution C 1*s*, O 1*s*, N 1*s* and S 2*p* spectra for p-CNO and the three CNO-PAMAM hybrids are reported in **Figure 5**. The parental p-CNOs show no surface oxygen, as the O 1*s* signal is completely absent on them whereas it is present in the three functionalized materials due to the presence of the amide and ester C=O carbonyl groups. For the nitrogen, it is noteworthy that the XPS analysis allows to distinguish two different contributions for the N 1*s* signal: one is due to the tertiary amide and amino groups at about 400 eV present in all the materials, but as the only contribution for **CNO-PAMAM 2.5G**; the second for the terminal $-NH_2$ group at about 402 eV, which is found in **CNO-PAMAM 2.0G** and **CNO-PAMAM 3.0G**. The percentage ratio, obtained by deconvolution of the N 1*s* peak, is in excellent accordance with the theoretical ratio of the different nitrogen atoms (for the 400 eV peak: **CNO-PAMAM 2.0G** 72% vs 69%; **CNO-PAMAM 3.0G** 84% vs 72%).

Table 2. XPS data of CNO-PAMAM *x*G hybrid materials. Relative atomic % are italicized.

	C 1*s*	O 1*s*	N 1*s*	S 2$p_{3/2}$	N/S	N/O
p-CNO	*100.0*[a] 284.5	– –	– –	– –	–	–
CNO-PAMAM 2.0G	*87.5* 284.5 (80%) 286.2 (15%) 288.0 (5%)	*5.5* 531.7 (76%) 533.3 (24%)	*6.2* 400.0 (72%) 401.8 (28%)	*0.8* 163.9 (56%) 168.2 (44%)	6.8	1.1
CNO-PAMAM 2.5G(I)	*86.4* 284.4 (68%) 285.0 (17%) 286.5 (12%) 288.6 (3%)	*6.9* 531.8 (64%) 533.8 (36%)	*5.8* 400.0 (100%)	*0.9* 163.8 (67%) 168.6 (33%)	6.3	0.8
CNO-PAMAM 3.0G(II)	*85.4* 284.5 (33%) 286.0 (53%) 288.6 (6%) 290.6 (8%)	*5.3* 532.4 (79%) 533.6 (21%)	*8.8* 400.7 (84%) 402.6 (16%)	*0.4* 164.4 (73%) 169.0 (27%)	12.8	1.7

Of extreme importance is the presence of the contribution of sulfur, given its critical bridging role in the disulfide-linked CNO-PAMAM *x*G hybrids. Indeed, the sulfur contribution in the XPS affirms the successful coupling of the PAMAM dendron core onto the CNO surface. More specifically, the observed S 2*p* contribution comprises two spin-orbit doublets, indicating the presence of two different species having the two spin-orbit coupling components, namely S 2$p_{3/2}$ and S 2$p_{1/2}$. The S 2$p_{3/2}$ component is centered at about 164 eV, attributable to thioether sulfur bound to a graphitic skeleton [56, 57], and the S 2$p_{1/2}$ component, centered at about 168.5 eV, occurs due to sulfonic sulfur [58, 59]. This finding is analogous to the addition of disulfides to other CNMs, wherein, during functionalization, the sulfur undergoes a partial oxidation to its sulphonic form [42, 43]. These results indicate that, in addition to the CNO–S–dendrimer bridge, the dendritic skeletons are also bound to the CNO surface by means of a CNO–SO_2–dendrimer bridge.

Finally, it is worth noting that, in the three materials, the atomic N/S ratio is almost identical for **CNO-PAMAM 2.0G** and **CNO-PAMAM 2.5G(I)** and almost double for **CNO-PAMAM 3.0G(II)**. Although these values are not exactly the real ratio, they perfectly reflect that **CNO-PAMAM 3.0G(II)** has an N/S ratio almost double that of the other two materials.

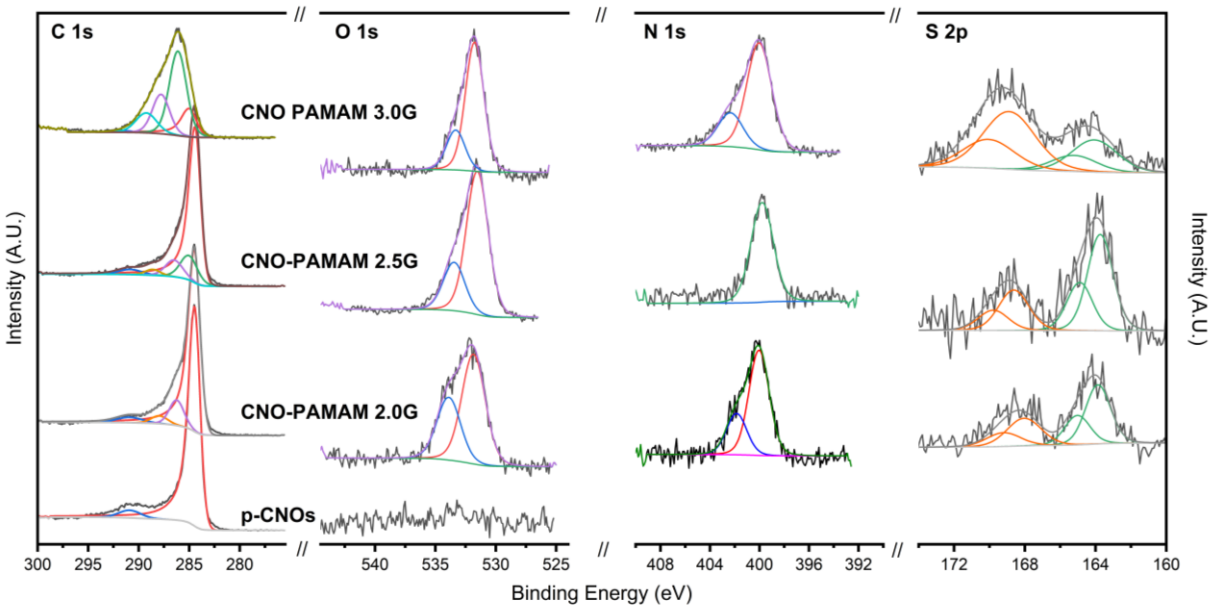


Figure 5. High-resolution XPS spectra (C 1*s*, O 1*s*, N 1*s*, S 2*p*) of p-CNOs and the three CNO-PAMAM *x*G materials.

Dynamic Light Scattering (DLS) and Zeta Potential (ZP) are two complementary techniques in the detailed characterization of carbon nanoparticles (CNPs). DLS analyses scattered light to infer the size distribution and aggregation behavior of nano- and microparticles in suspension. On the other hand, ZP reveals the surface charge inherent in these particles, which plays a crucial role in predicting their stability and interaction behavior in colloidal systems [60].

The analysis of these properties in CNPs is of significant interest, given the increasing drive for their use in nanopharmaceuticals and targeted drug delivery systems [31, 61, 62]. The aqueous solubility dispersion of CNPs remains a significant challenge for their transition into bio-applications, given the propensity for carbon nanoforms to aggregate [63]. To date, this challenge has been addressed through various surface modification approaches [64]. Consequently, we are led to an intriguing inquiry: can PAMAM enhance the solubility of CNOs in water-based solutions?

DLS and ZP (**Figure 6**) analyses of **CNO-PAMAM 2.0G**, **CNO-PAMAM 2.5G(I)** and **CNO-PAMAM 3.0G(II)** were carried out in deionized water at concentrations of 5, 10 and 20 ppm. For each sample, several measurements were taken, which are represented as faint traces in each subplot.

From these individual measurements, a representative average curve was calculated and visualized for each material at each concentration. This curve, shown as the thick line, represents the most common particle size in the suspension in DLS studies, and the most common particle surface potential in ZP studies. These DLS results, coupled with the ZP findings, provide crucial insights into the aqueous stability afforded by PAMAM-functionalization on CNOs.

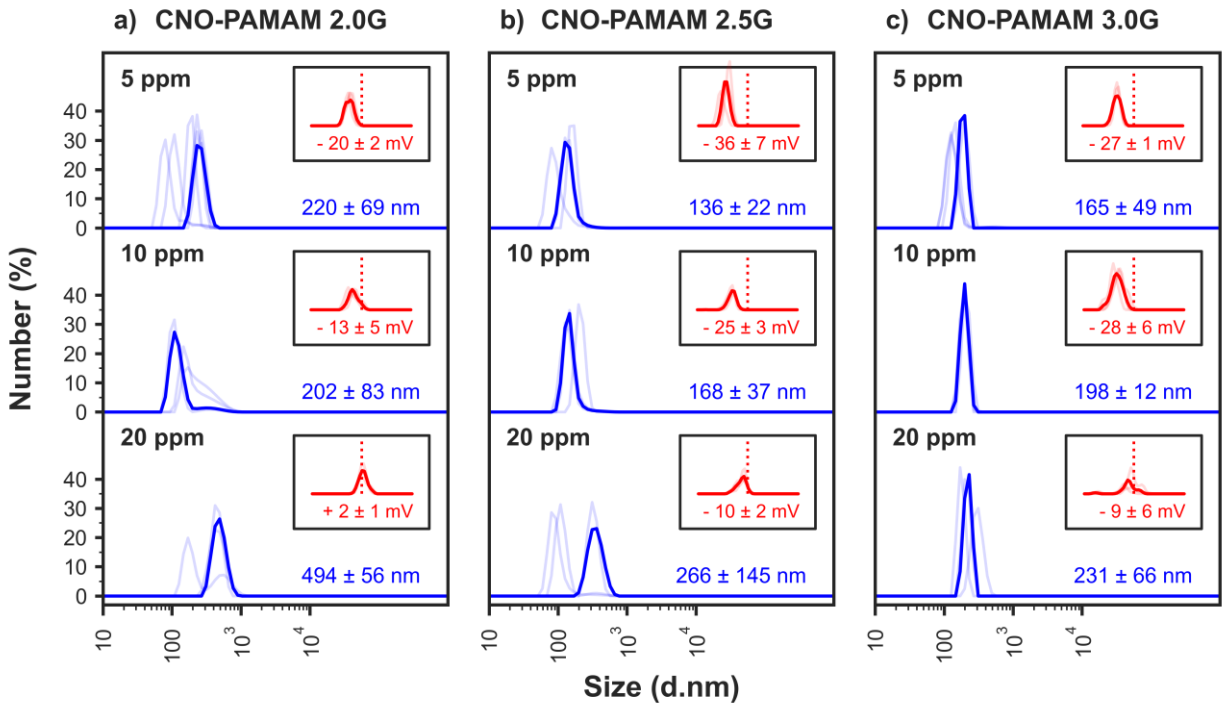


Figure 6. DLS plots (blue) with ZP inset plots (red) of three generations of **CNO-PAMAM *x*G** dispersed in deionised water at concentrations of 5, 10 and 20 ppm. The mean ± standard deviation hydrodynamic diameter and ZP values are noted ($n \geq 6$ for 5 ppm and $n \geq 3$ for 10 and 20 ppm). Individual measurements are shown as faint traces.

ZP results demonstrated optimal values, specifically less than -25 mV, when all materials were dispersed at 5 and 10 ppm concentrations. However, a ZP shift towards 0 mV was observed when the materials were dispersed at 20 ppm, indicating a greater propensity towards aggregation at this concentration. This indicates that the stability of CNO-PAMAM nanoparticles in water diminishes at concentrations exceeding 20 ppm. Of the three generations, **CNO-PAMAM 2.5G(I)** and **CNO-PAMAM 3.0G(II)** demonstrated superior ZP results. The ZP of **CNO-PAMAM 2.0G** was acceptable only at 5 ppm, with a value of approximately -20 ± 2 mV. At concentrations of 10 and

20 ppm, the ZP values for **CNO-PAMAM 2.0G** were not conducive to stability. This instability is corroborated by the DLS findings, which show a significant increase in particle size for **CNO-PAMAM 2.0G** at 20 ppm (494 ± 56 nm), and high deviation between individual measurements at 5 and 10 ppm. Hence, it can be inferred that the 2.0 generation of PAMAM dendrimers is suboptimal for inferring good aqueous dispersion on CNOs.

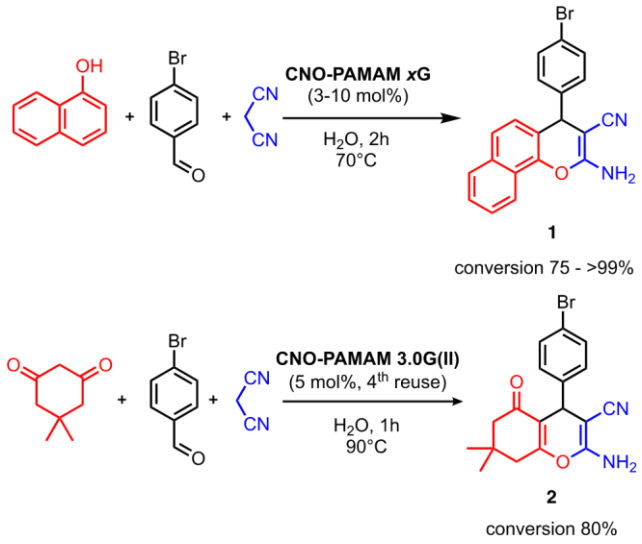
Notwithstanding, both **CNO-PAMAM 2.5G(I)** and **CNO-PAMAM 3.0G(II)** demonstrated promising performance, maintaining favorable average particle sizes (< 200 nm) at 5 and 10 ppm. Notably, **CNO-PAMAM 3.0G(II)** exhibited the most monodispersed distribution, as evidenced by its narrow distribution bands. The standout result was observed for **CNO-PAMAM 3.0G(II)** at 10 ppm, which consistently maintained an excellent distribution size of 198 ± 12 nm and a ZP of -28 ± 6 mV over multiple measurements.

The collective interpretation of these DLS and ZP results offers insights into the influence of PAMAM-functionalization on the aqueous stability of CNOs. Given the favorable results demonstrated by **CNO-PAMAM 2.5G(I)** and **CNO-PAMAM 3.0G(II)**, further studies can be envisaged that would contribute to the bio-applications of PAMAM dendrimers in improving CNOs dispersibility—namely, given the primary amino tail-end terminations of 3.0G, which are not present in the 2.5G PAMAM dendrimer, it can be speculated that a further improved dispersibility may be attained by lowering the pH of the aqueous dispersions.

To check the superficial availability of PAMAM amino groups, we decided to test **CNO-PAMAM 2.0G** and **CNO-PAMAM 3.0G(II)** as possible heterogeneous organocatalysts in the one-pot three-component synthesis in water of some derivative of 2-amino-4*H*-chromene (**Scheme 2**).

Functionalized 4*H*-chromenes are vital compounds because of their widespread presence in nature and biological activity [65], find therapeutic applications in medicinal chemistry as anti-HIV, anti-tumor, anti-inflammatory, antifungal, antimicrobial, and anti-allergenic [66-69] among others.

The one-pot synthesis of 2-amino-3-cyano-4-(4-bromophenyl)-4*H*-benzo[*h*]chromene **1** has been selected as the benchmark reaction, and a quick solvent screening prepositioned water as the best reaction media. The hybrids **CNO-PAMAM 2.0G** and **CNO-PAMAM 3.0G(I)** and **(II)** were first tested in 3 mol% content (referred to -NH_2 groups, entries 1-3), being the first the most active with a conversion of 89% (**Table 3**, entry 1). Nevertheless, the higher loading in amino-groups of **CNO-PAMAM 3.0G(II)** makes this catalyst more useful since just a few milligrams are needed to run the reaction, and a second attempt with 10 mol% of this hybrid gave complete formation of the 4*H*-chromene derivative (**Table 3**, entry 4).



Scheme 2. One-pot synthesis of 2-amino-4*H*-chromene derivatives mediated by CNO-PAMAM *x*G hybrids.

Table 3. CNO-PAMAM mediated synthesis of 2-amino-4*H*-chromene derivatives **1-2**.[a]

Entry	Catalyst	Loading (mol%)	Product	Conversion (%)
1	**CNO-PAMAM 2.0G**	3	**1**	89
2	**CNO-PAMAM 3.0G(I)**	3	**1**	75
3	**CNO-PAMAM 3.0G(II)**	3	**1**	81
4	**CNO-PAMAM 3.0G(II)**	10	**1**	>99
5[b]	**CNO-PAMAM 3.0G(II)**	5	**2**	80

[a.] Reaction conditions: 4-bromobenzaldehyde (0.2 mmol), 1-naphthol (0.2 mmol), malononitrile (0.22 mmol), catalyst, water (0.5 mL), 2 h, 70 °C; [b.] Reaction conditions: 4-bromobenzaldehyde (0.25 mmol), dimedone (0.25 mmol), malononitrile (0.3 mmol), catalyst, water (0.5 mL), 1 h, 90 °C.

The convenience of applying heterogeneous catalysis relies on the ease of catalyst recovery and its reuse. In this regard, the recyclability of **CNO-PAMAM 2.0G** (3 mol%) and **CNO-PAMAM 3.0G(II)** (10 mol%) was checked for the title reaction (**Figure 7**). In the case of **CNO-PAMAM 2.0G**, the conversion of the benzaldehyde into the chromene-derivative gradually diminishes from 89% to 70%. Nevertheless, this finding is not associated with loss in activity but with little losses of materials after the recovery, given that the turnover number (TON) associated with the actual catalytic loading stays in the 26-28 range between the second and the fifth cycle (**Figure 7**). On the other hand, catalyst **CNO-PAMAM 3.0G** yields almost quantitative substrate conversion within two hours in four consecutive runs (**Figure 7**).

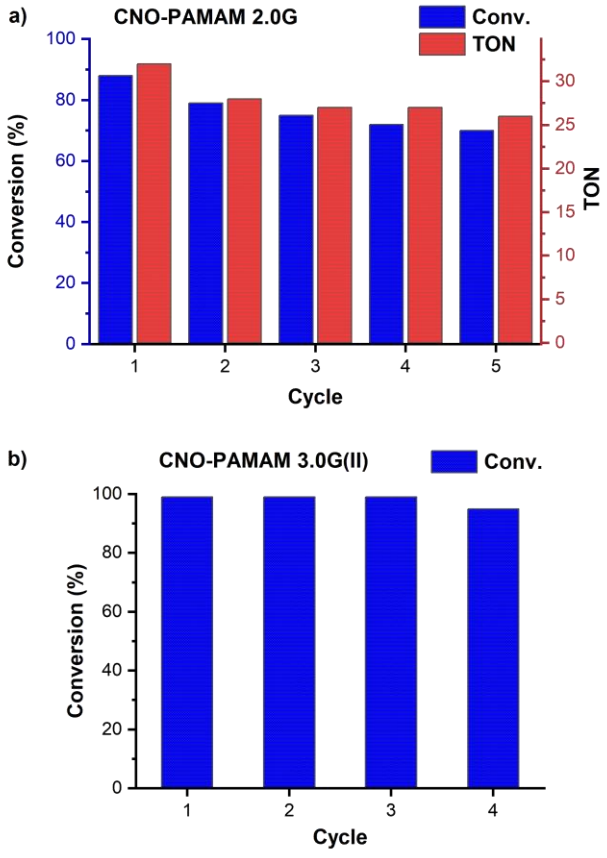


Figure 7. Recycles with (a) **CNO-PAMAM 2.0G** (3 mol%), and (b) **CNO-PAMAM 3.0G(II)** (10 mol%).

Finally, a reused **CNO-PAMAM 3.0G(II)** catalyst successfully converted 4-bromobenzaldehyde to the corresponding 2-amino-4*H*-chromene derivative in 80% yield when dimedone was used in place of 1-naphthol (**Scheme 2** and Table 3, entry 5).

4. Conclusion

A new easy and atom-economy protocol for the covalent functionalization of carbon nano-onions (CNOs) through the radical addition reaction of disulfides has been reported. Three different poly(amidoamine) (PAMAM) dendrimers of increasing generation were successfully bound onto the CNO surface, with corresponding hybrids indicating that reaction time imparts control over the degree of functionalization. These CNO-PAMAM hybrids were extensively characterized using various analytical and spectroscopic techniques.

The TGA was used to establish the degree of functionalization of CNO both in dendron and in amino groups (where present), indicating a loading as high as 3.48 mmol/g for **CNO-PAMAM 3.0G(II)**.

Useful information was obtained from the XPS analysis, which made it possible to establish without a doubt that the dendrimeric skeletons were actually bonded on the surface of the CNOs and that this occurs by means of CNO–S–dendrimer and CNO–SO_2–dendrimer bonds, being the sulfur partly oxidized during the radical addition process. In a key result of this study, for the first time it was possible to acquire solid-state CP-MAS ^{13}C-NMR spectra of CNO derivatives. The analysis of the DLS and Zeta potential data on the hybrids allowed to establish the dimensions of the aggregates and their stability in aqueous solutions, highlighting the goodness of these in the cases of **CNO-PAMAM 2.5G** and **CNO-PAMAM 3.0G**, which become thus excellent candidates for subsequent studies in the bio-application field.

Regarding their catalytic performance, **CNO-PAMAM 2.0G** and **CNO-PAMAM 3.0G** were found to be good heterogeneous catalysts in three-component one-pot processes for the synthesis of 2-amino-4*H*-chromene derivatives, resulting in easily recoverable and recyclable, with no losses in catalytic activity observed for at least five cycles.

These findings pave the way toward the preparation of other f-CNOs with different functional disulfides, for the preparation of more complex CNO-based catalytic systems, and for the preparation of new hybrid systems for drug delivery or with biological activity.

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